1. The following common acids are strong: HCl, HBr, HI, HNO\textsubscript{3}, HClO\textsubscript{4}, H\textsubscript{2}SO\textsubscript{4}

   The following are some less common acids that are also strong: HClO\textsubscript{3}, HBrO\textsubscript{3}, HIO\textsubscript{3}, H\textsubscript{2}SeO\textsubscript{4}

   Assume all other acids are weak unless told otherwise.

   Some weak acids: HF, HNO\textsubscript{2}, HClO\textsubscript{2}, [H\textsubscript{2}SO\textsubscript{3}] = SO\textsubscript{2} + H\textsubscript{2}O, HC\textsubscript{2}H\textsubscript{3}O\textsubscript{2} = HOAc

2. All ionic hydroxides are strong bases, regardless of solubility. Bases that do not contain OH\textsuperscript{−} are weak.

   Some strong bases: LiOH, NaOH, KOH, RbOH, CsOH, Mg(OH)\textsubscript{2}, Ca(OH)\textsubscript{2}, Sr(OH)\textsubscript{2}, Ba(OH)\textsubscript{2}

   Some weak bases: NH\textsubscript{3}, CH\textsubscript{3}NH\textsubscript{2}, C\textsubscript{2}H\textsubscript{5}NH\textsubscript{2}

   Not a base: CH\textsubscript{3}OH, C\textsubscript{2}H\textsubscript{5}OH